

Additions and Corrections

Contribution of Tunneling to Secondary Isotope Effects in Proton-Transfer Reactions [*J. Am. Chem. Soc.* **1984**, *106*, 2223–2224]. WILLIAM H. SAUNDERS, JR.

Page 2224, eq 5 should read: $1 - A^2 - 2B^2 - 3C^2 + A^2B^2 + 3B^2C^2 = D$.

Page 2224, paragraph 2: The effect of the above change is to make D vary within a given model for models 3–5. Model 3: $D = -0.21$ to -0.31 . Model 4: $D = -0.32$ to -0.68 . Model 5: $D = -0.37$ to -0.87 . Models based on the correct equation given somewhat different numerical values, but none of the qualitative conclusions is changed.

Dimethylsilylene: Its Optical Absorption Spectrum and Reaction Kinetics [*J. Am. Chem. Soc.* **1984**, *106*, 7267–7268]. A. S. NAZRAN, J. A. HAWARI, D. GRILLER,* I. S. ALNAIMI, and W. P. WEBER*

The reported optical absorption spectrum assigned to dimethylsilylene ($\lambda_{\max} = 350$ nm) was probably due to photolysis of trace impurities contained in samples of its precursor, dodecamethylcyclohexasilane. When carefully purified by column chromatography, this material showed only a very weak absorption at the photolysis wavelength used in the experiments (308 nm). Likely impurities which have extremely intense absorption bands in this region are linear oligomers or polymers of dimethylsilylene. In view of these uncertainties we wish to withdraw our assignment until a reinvestigation is complete.

Heavy-Atom Kinetic Isotope Effects in the Acid-Catalyzed Rearrangement of *N*-2-Naphthyl-*N'*-phenylhydrazine. Rearrangement Is Shown To Be a Concerted Process [*J. Am. Chem. Soc.* **1985**, *107*, 6674–6678]. HENRY J. SHINE,* LIDIA KUPCZYK-SUBOTKOWSKA, and WITOLD SUBOTKOWSKI

Two errors were made in the treatment of raw data in the calculations of carbon KIE. These errors were in the treatment of "del" values in the ^{13}C isotope-ratio data and in omitting intramolecular competition in the ^{13}C and ^{14}C calculations. Re-treatment of "del" data and calculating KIE on the (only available) basis that intra- and intermolecular KIE in the rearrangement of the title compound (**1**) are of equal magnitude gives corrected results for the sought-after intermolecular KIE. The averaged results given in the abstract should be 1.0176 ± 0.0003 for one ^{13}C atom and 1.0287 ± 0.0011 for one ^{14}C atom. These corrections should also appear in the Discussion section of p 6677. The ^{14}C KIE calculated from eq 7 should read (line 32, p 6677) 1.0150 as compared with the experimental result (line 32) 1.0176. The last five entries in column 6, Table II should read in descending order as follows: 1.0252, 1.0259, 1.0281, 0.0174, 1.0162. Similarly, the last five entries in column 7 should read as follows: 1.0284, 1.0276, 1.0302, 1.0178, 1.0173. The transition state for rearrangement of **1** thus, contrary to the earlier indication, is unlike that in the rearrangement of 2,2'-hydrazonaphthalene and more like that in the rearrangement of hydrazobenzene.

Preparation and Characterization of Stable Ruthenium(IV)-Oxo Complexes Which Contain Tertiary Phosphine Ligands [*J. Am. Chem. Soc.* **1986**, *108*, 510–511]. MARY E. MARMION and KENNETH J. TAKEUCHI*

Page 511, left column, final paragraph, fifth sentence: The reported rate constants are incorrect. The sentence should read:

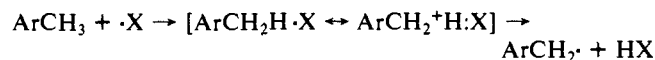
With $[(\text{bpy})_2\text{Ru}^{\text{IV}}(\text{O})(\text{PPh}_3)]^{2+}$ as the catalyst, the oxidation of benzyl alcohol has a second-order rate constant of $1.05 \text{ M}^{-1} \text{ s}^{-1}$, while the reaction involving $[(\text{bpy})_2\text{Ru}^{\text{IV}}(\text{O})(\text{PET}_3)]^{2+}$ has a second-order rate constant of $4.2 \times 10^{-3} \text{ M}^{-1} \text{ s}^{-1}$.

Nucleophilic Attack of Amine and Hydroxide to Platinum Dibenzonitrile Dichloride. Crystal Structure of $[\text{Pt}(\text{NH}=\text{CPhN}-t\text{-BuCH}_2\text{CH}_2\text{NH}-t\text{-Bu})\text{Cl}(\text{NHCOPh})]$ (2**) and *cis*- $[\text{Pt}(\text{NH}=\text{CPhN}-t\text{-BuCH}_2\text{CH}_2\text{NH}-t\text{-Bu})\text{Cl}(\text{NCPH})]$ (**3**)** [*J. Am. Chem. Soc.* **1986**, *108*, 1180]. LUCIANA MARESCA, GIOVANNI NATILE,* FRANCESCO P. INTINI, FRANCESCO GASPARRINI, ANTONIO TIRIPICCHIO, and MARISA TIRIPICCHIO-CAMELLINI

Page 1181, second full paragraph, 4th row before the end should read: ...absorption at 2280 assignable to $\nu_{\text{C}\equiv\text{N}}$; other significant bands are at 1565 ($\nu_{\text{C}=\text{N}}$), ...

Temperature Dependence of the Homolytic Abstraction of Benzylic Hydrogen by Bromine Atom [*J. Am. Chem. Soc.* **1986**, *108*, 2340]. J. G. WRIGHT and R. D. GILLIOM*

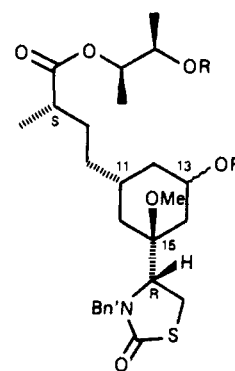
Page 2340: The equation shown should be



Page 2341: The legend for Figure 2 should be the following: Hammett plot for the relative rates of substituted toluenes, k' , and toluene, k , at -50°C .

Total Synthesis of (+)-Latrunculin B [*J. Am. Chem. Soc.* **1986**, *108*, 2451–2453]. REGINA ZIBUCK, NIGEL J. LIVERTON, and AMOS B. SMITH III*

Page 2452: structures 16a and 16b are incorrect and should be as illustrated below:



16a R = H

16b R = TBS

Bn' = $p\text{-MeOC}_6\text{H}_4\text{CH}_2\cdot$

New Synthetic Technology for the Construction of Oxocenes [*J. Am. Chem. Soc.* **1986**, *108*, 2468–2469]. K. C. NICOLAOU,* M. E. DUGGAN, and C.-K. HWANG

Page 2469: footnote 16, should read 200 MHz and not 50.3 MHz; footnote 18 should read ...axial (α) attack... and not ...axial (β) attack...